

ORIGINAL RESEARCH PAPER

Fabrication Fe₃O₄/SiO₂/TiO₂ Nanocomposites and Degradation of Rhodamine B Dyes under UV Light Irradiation

M. P. Mazhari*¹, A. Abbasi¹, A. Derakhshan², M. Ahmadi²

¹Researchers and Elite club, Marivan Branch, Islamic Azad University, Marivan, Iran

²Department of Chemistry, Catalysis Division, University of Isfahan, Isfahan

ARTICLE INFO.

Received 12/10/2015

Accepted 23/12/2015

Published online 01/01/2016

KEYWORDS

Magnetic nanoparticle

Nanocomposite

photocatalyst

Sol-gel

ABSTRACT

Recycling and reusing of catalyst is an important factor to produce capable and low cost catalysts. Silica coated magnetic nanoparticles (Fe₃O₄/SiO₂) were synthesized via a simple sol-gel method with the aid of sonication. Fe₃O₄ nanoparticles. After that a layer of TiO₂ was constructed by hydrolyze and condensation of Teteranormalbuthyltitanate to produce Fe₃O₄/SiO₂/TiO₂ nanocompsite. As-synthesized nanoparticles were characterized with X-Ray powder Diffraction, Fourier-Transform-Infrarotspektrometer, Transition Electron Microscopy and Scanning Electron Microscopy. In this study, a new kind of Fe₃O₄/SiO₂/TiO₂ photocatalyst was prepared with the purpose of using light and inhibiting the recombination of electrons and holes. In situ treatment of Rhodamine B in water was performed using this Fe₃O₄/SiO₂/TiO₂ photocatalyst. The purpose of this research was to provide a new method and the basic data for the removal of organic pollutants in water. Fe₃O₄/SiO₂/TiO₂ nanocomposite showed enhanced photocatalytic properties as evidenced by the enhanced photodegradation of Rhodamine B under ultra violet light irradiation.

INTRODUCTION

Magnetic nanoparticles are currently being investigated because of their numerous applications in several fields such as, medicine, magnetic resonance imaging, data storage, , and water remediation . It is often necessary to coat their surface with an organic or inorganic shell, in order to protect them from chemical degradation or agglomeration according to the environments in which they will be used.

The coating can also be performed in order to add new functionalities to the magnetic core, such as biological stealth, optical properties, catalytic or adsorbing capacity.

Over the past decades, titanium dioxide (TiO₂) nanoparticles have gained much attention as a photocatalyst and catalyst support [1]. TiO₂

nanoparticles have many advantages as compared to other photocatalysts, which include excellent high stability against chemical and photonic corrosion and high photocatalytic activity [2].

TiO₂ nanoparticles of small mean particle sizes possess high surface area and photocatalytic activity. However, TiO₂ nanoparticles of high surface area are thermally unstable and lose their surface area readily . Therefore, much effort has been focused on coating of TiO₂ on high surface area supports such as silica or alumina in order to stabilize TiO₂ nanoparticles. TiO₂ nanoparticles could be difficult to recover and lost readily upon being dispersed into wastewater. One of the ways to overcome this problem is to coat TiO₂ onto magnetite (Fe₃O₄) cores and the resulting Fe₃O₄/TiO₂ Nanocomposite nanoparticles can be recovered easily through manipulation by external magnetic field. However, it was difficult to achieve complete coating

✉ *Corresponding author

Email address: sampleEmail@yahoo.com

Tel.: +98 9189760059

of Fe_3O_4 nanoparticles with TiO_2 at nanometer scale using the sol-gel method [2]. Besides, TiO_2 would oxidize Fe_3O_4 nanoparticles and lead to a reduction of magnetic moment. Some researchers had attempted to coat a thin layer of SiO_2 between Fe_3O_4 nanoparticles and TiO_2 shell. The presence of a SiO_2 layer between TiO_2 shell and Fe_3O_4 nanoparticles could increase the lifetime of photogenerated holes which in turn, resulted in increased photoreactivity [3]. This is attributed to the SiO_2 layer which serves as an insulating layer between Fe_3O_4 nanoparticles (hole-electron trap center) and the TiO_2 shell. Herein, we have reported a facile and efficient synthesis approach for the fabrication of $\text{Fe}_3\text{O}_4/\text{SiO}_2/\text{TiO}_2$ discrete Nanocomposite nanoparticles by the sol-gel method with the aid of sonication [4]. Fe_3O_4 nanoparticles were being encapsulated inside discrete SiO_2 nanospheres within 90 minutes, and a TiO_2 layer was then coated directly onto each SiO_2 nanosphere via the sol-gel method. The photocatalyst properties of as-synthesized $\text{Fe}_3\text{O}_4/\text{SiO}_2/\text{TiO}_2$ Nanocomposite were evaluated by the photodegradation of Rhodamine B with or without UV light irradiation [5,6].

MATERIALS AND METHODS

Reactions of catalyst preparing were done under an inert argon or nitrogen atmosphere. Polyethylene glycol 400 (PEG 400), hydrochloric acid, ammonium hydroxide 30%, 2-propanol, iron(II) chloride hydrate, iron(III) chloride hydrate, tetraethyl orthosilicate (TEOS), tetranormalbuthyltitanate (TNBT), acetylacetone, ethanol and methanol were purchased from Merck company and all of them used without further purification. During all steps when water was needed, deionized water was used throughout.

FT Infrared (FT-IR) spectra were obtained as potassium bromide pellets in the range of 400–4000 cm^{-1} with a 6300 JASCO Perkin Elmer spectrophotometer. XRD patterns were recorded by a Bruker D8 Advanced, X-ray diffractometer using Ni-filtered $\text{Cu K}\alpha$ radiation. SEM images were obtained using a VEGA-II TESCAN. TEM images were obtained using a CM120. The UV-vis spectra of the samples were taken on a UV-vis spectrophotometer (Shimadzu, UV-2550, Japan) + visible sources of 400 W Osram lamps.

Synthesis of the catalyst support ($\text{Fe}_3\text{O}_4/\text{SiO}_2$ nanoparticles)

The catalyst support comprised of silica-coated magnetic NPs was synthesized by a co-precipitation method. These nanoparticles were prepared by chemical co-precipitation of Fe^{3+} and Fe^{2+} ions with a molar ratio of 2:1. Typically, $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$ (5.838 g) and $\text{FeCl}_2 \cdot 4\text{H}_2\text{O}$ (2.147 g) were dissolved in 38 mL degassed HCl 0.4 molar. Then, this solution was added to 300 mL of degassed 25% NH_4OH quickly in one portion while stirring in ultrasound bath. The addition of the base to the $\text{Fe}^{2+}/\text{Fe}^{3+}$ salt solution resulted in the formation of the black precipitate of MNPs immediately. Stirring in ultrasound bath continued for another 90 min. Black solid collected by magnet and washed with ethanol and distilled water three times. Then the sediment dispersed in 150 ml distilled water.

Since Fe_3O_4 nanoparticles are delicate to heat and oxidation milieu, to protect their surface and to prevent oxidation to Fe_2O_3 , surface of magnetic nanoparticles coated by a thin layer of SiO_2 . The SiO_2 shell can effectively prevent chemical degradation and photodissolution property of Fe_3O_4 particles making the photocatalyst recyclable after multiple reaction cycles, and promote the photocatalytic activity of TiO_2 by decreasing the adverse influence of magnetic core. At the second step for synthesis of silica-coated Fe_3O_4 nanoparticles, 50 mL of previous suspension added to a solution of 250 mL 2-Propanol that contained ammonium hydroxide (4.5 mL). This dispersion was homogenized by ultrasonic vibration. Then, 6 mL tetraethoxysilane (TEOS) was slowly added. Finally, silica was formed on the surface of magnetite nanoparticles through hydrolysis and condensation process. The resulting silica-coated Fe_3O_4 nanoparticles were thoroughly washed with deionized water and collected by magnetic separation, followed by drying at 45 °C under vacuum for 12 h.

Synthesis of $\text{Fe}_3\text{O}_4/\text{SiO}_2/\text{TiO}_2$ nanocomposite

Specifically, 0.7 g of $\text{Fe}_3\text{O}_4/\text{SiO}_2$ nanoparticles were dispersed in 150 ml of merck ethanol and sonicated for 20 minutes. After that 5 ml PEG 400 was added to the mixture. In second vessel, 8 ml Tetranormalbuthyltitanate was added to 20 ml of Merck ethanol and 1 ml acetylacetone and stirred on magnetic stirrer for 10 minutes. Second vessel was added slowly to above mixture while gently shaking by mechanical stirrer. After 20 minutes, 3 ml deionized water was added and final mixture stirred at 70 °C for 12 hours. The obtained gray products were separated by an external magnet and washed thoroughly with ethanol and

deionized water before drying at 60 °C. Subsequently, the powder was calcined at 400°C in air for 2 h to improve the crystallinity of the obtained Fe₃O₄@SiO₂@TiO₂ nanoparticles.

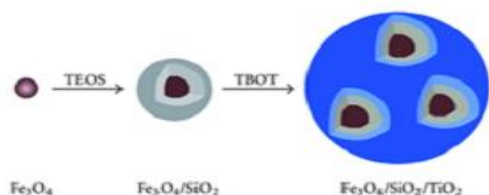


Fig. 1. Procedure for the synthesis of Fe₃O₄@SiO₂@TiO₂

RESULTS AND DISCUSSION

The morphology of Fe₃O₄@SiO₂@TiO₂ samples powder were evaluated by scanning electron microscopy (SEM) illustrated in (Fig. 2). Images reveals that for Fe₃O₄@SiO₂@TiO₂ synthesized by coprecipitation method that has been calcinated in 400 °C, particles almost are monosize with mediocre size about 30 nm.

Fe₃O₄ nanoparticles were being encapsulated within the SiO₂ shells upon the hydrolysis and condensation of TEOS as new bonds of Fe–O–Si were formed between the interface of Fe₃O₄ and TEOS. Ultrasonication was used to accelerate the hydrolysis of TEOS. This was followed by lateral polymerization, and the formation of a three-dimensional network via siloxane formation (Si–O–Si), to produce a homogenous SiO₂ coating.

(SiO₂) between TiO₂ coating and magnetic material was proposed for better coating of TiO₂ layer and to avoid photodissolution of iron.

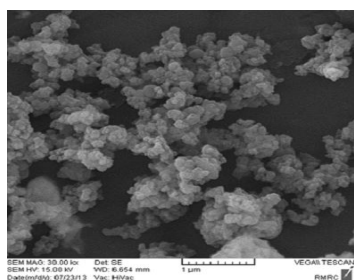


Fig. 2. SEM image of Fe₃O₄/SiO₂ (a) Fe₃O₄/SiO₂/TiO₂ magnetic nanoparticle.

The synthesized nanocomposite was characterized using FT-IR (Fig. 3). The background corrected FT-IR

spectra of MNPs are shown in (Fig. 3. a). The broad band around 3400 cm⁻¹ can be assigned to O-H stretching vibration which is assigned to the surface OH groups of Fe₃O₄ NPs. The absorption bands around 580-610 cm⁻¹ is attributed to the vibration of Fe-O bond in Fe₃O₄. In the FT-IR spectrum of Fe₃O₄@SiO₂ and Fe₃O₄@SiO₂/TiO₂, the band at about 1100 cm⁻¹ is assigned to stretch of Si-O bond and the SiO-H/TiO-H groups are appeared by the very broad IR absorption band in the 2800-3700 cm⁻¹ region. In the (Fig. 3c), the ratio of 580-610 cm⁻¹ to the 1100 cm⁻¹ show that the coating of TiO₂ is not core-shell nanoparticle. The structure was synthesized as nanocomposite format (Fe₃O₄@SiO₂/TiO₂) [7].

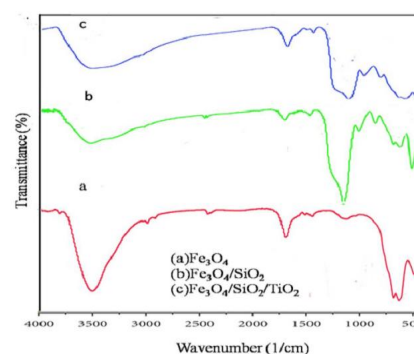


Fig. 3. FT-IR a) Fe₃O₄ nanoparticles b) Fe₃O₄/SiO₂ c) Fe₃O₄/SiO₂/TiO₂

The XRD patterns of photocatalysts are compared in Fig. 4. Only the diffraction peaks of typical anatase TiO₂ were observed with 2θ at 25.3°, 37.9°, 48.2°, 54.0°, 55.1°, 62.6°. Diffraction peaks of titanate and Fe₃O₄ crystal didn't appear. This indicates that the Fe₃O₄ is well encapsulated by SiO₂, and no interaction between Fe₃O₄ and TiO₂ occurs [8,9].

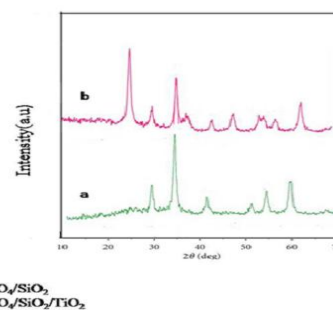


Fig. 4. XRD pattern of Fe₃O₄/SiO₂ (a) and Fe₃O₄/SiO₂/TiO₂ (b)

Fig. 5 shows Transmission electron microscopy (TEM) image of Fe₃O₄/SiO₂/TiO₂ nanoparticles. Based on TEM image, the ultrafine particle diameter size of the synthesized Fe₃O₄ is about 20-40 nm.

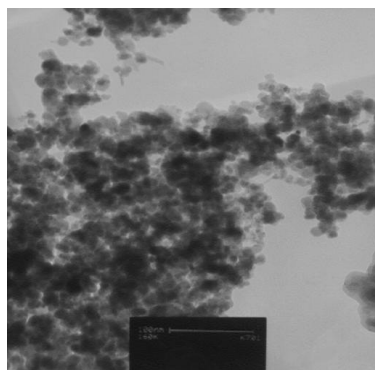


Fig4. TEM image of Fe₃O₄/SiO₂/TiO₂ magnetic nanoparticle

Photocatalyst degradation of Rhodamine B by Fe₃O₄@SiO₂@TiO₂ nanoparticle was performed under ultraviolet light irradiation. Changing in dye concentration was handled by UV absorption spectroscopy. Amount of the Rhodamine B degradation in time of t (DP(t)) was calculated as follows:

$$DP(t) = \frac{A_0 - A_t}{A_0} \times 100$$

Where A₀ and A_t are the absorbance value of the solution at 0 and t minute, respectively.

To proving catalysts effect, the mixture stirred in dark milie. No dye degradation after 40 minutes without using UV light irradiation or nanopowder photocatalysts, so the contribution of self-degradation was insignificant. As time passed, dye decomposed and the color of the solution became brighter until at t = 40 min a rather transparent solution was obtained. Results for the changes in the concentration of dye after 40 min under UV irradiation.. After 40 minutes, the degradation percentage of Fe₃O₄@SiO₂@TiO₂ - obtained 64% (Fig. 6).

CONCLUSION

In this study, a simple and facile synthesis approach was developed for the preparation of a magnetically separable photocatalyst consisting of an Fe₃O₄ core, an SiO₂ intermediate layer, and a photocatalytically active TiO₂. This synthesis method was rapid and did

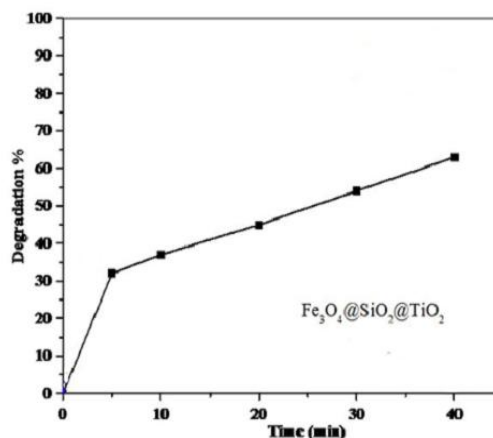


Fig. 6. Degradation of Rhodamine B dyes by Fe₃O₄/SiO₂/TiO₂ magnetic nanoparticle

not require the addition of any surfactant to direct the formation of SiO₂ or TiO₂ shells. The photocatalytic activity of TiO₂ surface shell was not affected by the intermediate SiO₂ layer and Fe₃O₄ core. The Fe₃O₄/SiO₂/TiO₂ nanoparticles possessed high specific surface area and exhibited a good photocatalytic activity for the photodegradation of Rhodamine B dye in aqueous solution

CONFLICT OF INTEREST

The authors declare that there are no conflicts of interest regarding the publication of this manuscript.

REFERENCES

- [1] E. Beyers, E. Biermans, S. Ribbens, K. De Witte, M. Mertens, V. Meynen, S. Bals, G. Van Tendeloo, E. F. Vansant, P. Cool, *Applied Catalysis B: Environmental*. 88 (2009), 515-524.
- [2] S. C. Pang, S. Y. Kho, S. F. Chin, *Journal of Nanomaterials*. (2012).
- [3] T. A. Gad-Allah, S. Kato, S. Satokawa, T. Kojima, *Solid State Sciences*. 9 (2007), 737-743.
- [4] A.-L. Morel, S. I. Nikitenko, K. Gionnet, A. Wattiaux, J. Lai-Kee-Him, C. Labrugere, B. Chevalier, G. Deleris, C. Petibois, A. Brisson, M. Simonoff, *ACS Nano*. 2 (2008), 847-856.
- [5] Osuna, K. M. Gregorio-Jauregui, J. G. Gaona-Lozano, I. M. d. I. Garza-Rodr, , guez, A. Ilyna, E. D, a. Barriga-Castro, H. Saade, Ra., G. L, , pez, *J. Nanomaterials*. 2012 (2012), 103-103.
- [6] S. C. Pang, S. Y. Kho, S. F. Chin, *J. Nanomaterials*. 2012 (2012), 125-125.
- [7] L. E. Ahangar, K. Movassaghi, M. Emadi, F. Yaghoobi, *Nano. Chem. Res*. 1 (2016), 33-39.
- [8] Z. Peng, H. Tang, Y. Tang, K. F. Yao, H. H. Shao, *International Journal of Photoenergy*. 2014 (2014), 8.

- [9] Y. Kobayashi, S. Saeki, M. Yoshida, D. Nagao, M. Konno, *Journal of Sol-Gel Science and Technology*. 45 (2007), 35-41.

AUTHOR(S) BIOSKETCHES

Mazhari, M.P., Ph.D., Ph.D., Research Scientist, Researchers and Elite club, Marivan Branch, Islamic Azad University, Marivan, Iran. Email: Peyman.Mazhari@yahoo.com

Abbasi, A., Ph.D., Research Scientist, Department of chemistry, University of Isfahan, Isfahan-52795, Iran. Email: Ali.Abbasi862@gmail.com

Derakhshan, A., Ph.D., Department of chemistry, University of Isfahan, Isfahan-52795, Iran. Email: Abbas.Derakhshan64@yahoo.com

Ahmadi, M., Ph.D., Research Scientist, Department of chemistry, University of Isfahan, Isfahan-52795, Iran. E-mail: M.Ahmadi67@gmail.com

How to cite this article:

Mazhari M.P, Abbasi A, Derakhshani A, Ahmadi M. Fabrication $Fe_3O_4/SiO_2/TiO_2$ Nanocomposites and Degradation of Rhodamine B Dyes under UV Light Irradiation. J. Nanostruct. 2016; 6(1):101-105.

DOI: 10.7508/jns.2016.01.017

URL: http://jns.kashanu.ac.ir/article_13698.html