

## Studies on Properties of Ag\Al<sub>2</sub>O<sub>3</sub> Nanocomposite Synthesized by Solution Combustion Technique Under Novel Condition

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### Abstract

Ag\Al<sub>2</sub>O<sub>3</sub> nanocomposite were synthesized by solution combustion Technique using ammonium carbonate as a new fuel under simultaneous combustion of alumina and silver particles condition. Samples prepared with different amount of silver, were characterized by XRD, BET, SEM and optical Absorption measurements. XRD analysis was used to determine the change in phase and size of the synthesized nanoparticle with adding silver. BET analysis showed that added silver reduced specific surface but which is still suitable for catalytic application. BJH analysis showed that pores in Ag/Al<sub>2</sub>O<sub>3</sub> nano composite are mesoporous and in Al<sub>2</sub>O<sub>3</sub> nano particles is microporous. TEM and SEM showed that silver was distributed completely homogenously in pores of Alumina, and didn't change in surface of Alumina. Optical transmission showed the different existence of different Ag species.

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## 1. Introduction

The functional nanocomposites involving ceramic matrix and nanosized particles of transition metals are known to exhibit multifunctional and attractive properties and are identified as potential candidate for structural [1], mechanical [2], catalytic [3], thermal [4], optical [5], magnetic [6] and electrical [7] applications. Nanocomposites are usually involved noble or transition metals incorporated in such oxide matrices such as alumina, silica and titania by

various processes whether as a catalyst or membrane [8, 9]. The nanocomposites of alumina with transition metals are reported to exhibit different and often improved properties compared to its bulk counterpart [10–11, 12]. Tuan et al. have shown that silver inclusions improve the mechanical [13] and transport properties [14] of alumina. At room temperature in air, flexural strength, fracture toughness, electrical resistivity, thermal diffusivity and electrical conductivity reach the following values when the silver content

is 10 vol.%: 394 MPa, 4.7 MPa m<sup>1/2</sup>, 11 × 10<sup>13</sup> Ω cm, 14.7 mm<sup>2</sup> s<sup>-1</sup>, and 40 W m<sup>-1</sup> K<sup>-1</sup> (compared to 325 MPa, 4.0 MPa m<sup>1/2</sup>, 1.8 × 10<sup>13</sup> Ω cm, 11.3 mm<sup>2</sup> s<sup>-1</sup>, and 32 W m<sup>-1</sup> K<sup>-1</sup> for pure alumina). Moreover, it was shown that these composites have an interfacial thermal conductance of 4.1 × 10<sup>8</sup> W m<sup>-2</sup> K<sup>-1</sup> [15].

Studies on the features of hardness in alumina silver composites are reported by Dutta et al., where the Knoop hardness values of the composites sensitively depend on the applied load of indentation [16]. Ag\Al<sub>2</sub>O<sub>3</sub> nanocomposites can be synthesized mainly by wet chemical [17], solvothermal- calcination (SC), impregnation-calcination(IC)[18], solgel, impregnation, incipient wetness impregnation [19] and alternating current[20] methods. However, all these techniques are quite involved and need high temperatures and long processing times. Moreover alumina and silver nanoparticles are synthesized and prepared separately but in present work we have synthesized both nano particles simultaneously using combustion process.

Solution combustion synthesis (SCS) is a versatile, simple and rapid process, which allows effective synthesis of a variety of nanosize materials [21].

In this study, we report the synthesis of Ag\g-Al<sub>2</sub>O<sub>3</sub> nanocomposite by solution combustion technique using a new type of fuel, Ammonium carbonate is a new mineral material and has lower price compared to often organic fuels and it has good potential to product nano materials with high surface area because of releasing vigorous enormous amount of gas while combustion process. The Ag\Al<sub>2</sub>O<sub>3</sub> nanocomposite prepared with different amount of Ag, have been characterized by different analysis such as XRD, BET, optical transmission and SEM. To assess the

change in properties of products with change in amount of silver.

## 2. Experimental procedure

Analytical grade of aluminium nitrate nanohydrate Al(NO<sub>3</sub>)<sub>3</sub>.9H<sub>2</sub>O(Merck) silver nitrate AgNO<sub>3</sub> and ammonium carbonate (NH<sub>4</sub>)<sub>2</sub>CO<sub>3</sub> (Merck) were used as the starting materials and ratio between aluminium nitrate and silver nitrate was chosen so that final Ag\Al<sub>2</sub>O<sub>3</sub> nano composite samples contained( 1wt%, 5wt% and 10wt% of Ag).

The starting materials were used including amount of stoichiometric of aluminium nitrate, silver nitrate to produce pure Al<sub>2</sub>O<sub>3</sub>(sample a), 1 wt% Ag\Al<sub>2</sub>O<sub>3</sub>(sample b), 5wt% Ag\Al<sub>2</sub>O<sub>3</sub>(sample c) and 10 wt% Ag\Al<sub>2</sub>O<sub>3</sub>(sample d) The preparation method for the Solution combustion synthesis was the same as that in the literature [22]. And then the solution put into ultrasound machine for 10 minute and 40C°, Finally The samples were calcined at 900 C° for 2 h.

## 3. Characterization

The crystallinity and phase identification of the powders were recorded by X-ray diffractometer (D4 Bruker) using Cu Kα as the radiation source and Ni as the filter over 2θ range from 10° to 80° at a rate of 2.5 °/ min. The voltage and the current setting were 30 kV and 30 mA, respectively. The samples were continuously scanned with a step size of 0.06° (2θ) and a count time of 1s per step. The strongest peaks in XRD patterns for Ag\Al<sub>2</sub>O<sub>3</sub> and alumina were selected to calculate the crystallite size by the Scherrer formula ( $d = 0.9\lambda/\beta\cos(\theta)$  where, d; crystal size, λ; wavelength of X-ray, θ; Bragg angle and β; peak broadening).

BET surface area and pore size distribution of synthesized powders were determined by nitrogen adsorption at 77 K using BET and BJH method respectively, by adsorption analyzer (BEL Japan, Inc). samples were observed by SEM(scanning electron Microscopy, Hitachi, japan). Transmission electron microscopy (TEM) was carried out with a CM300 FEG PHILIPS microscope to determine shape and average size of the particles. By using Diffuse-reflectance UV/Vis (spectra (UV/Vis) were recorded under ambient conditions on a U-3010(Hitachi spectrometer with a standard diffuse reflectance unite) had Study on the existence of different silver species into samples synthesis.

#### 4. Results and discussion

XRD pattern of pure  $\text{Al}_2\text{O}_3$  and  $\text{Ag}\backslash\text{Al}_2\text{O}_3$  nanocomposite (synthesized under the same condition) are shown in fig.1.

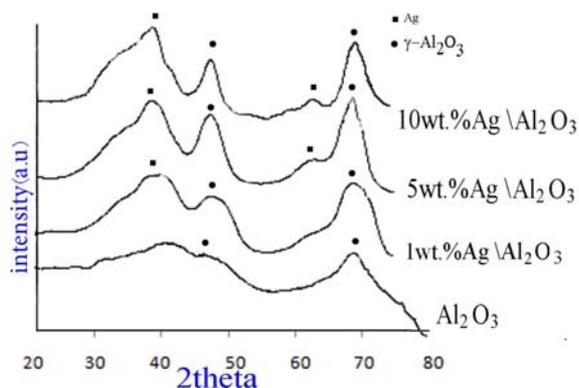


Fig. 1. XRD pattern of as- synthesized sample.

In all the samples the  $\gamma\text{-Al}_2\text{O}_3$  phase was detected. In XRD patterns of pure Alumina broad peak was observed at  $2\theta$  of 45 but this peak with added silver became much significant and sharper indicating the mean size of the  $\gamma\text{-Al}_2\text{O}_3$  particle in the  $\text{Ag}\backslash\text{Al}_2\text{O}_3$  nano composite increased. The Ag phases was observed at  $2\theta$  of 38.2 which correspond to the (111) lattice planes of metallic

Ag in the all samples, this might because that Ag species on the  $\text{Ag}\backslash\text{Al}_2\text{O}_3$  was crystalliy phase. It is observed that with increasing amount of Ag, the Ag (111) diffraction peak became much significant and sharper, indicating that the mean size of the Ag particle in the  $\text{Ag}\backslash\text{Al}_2\text{O}_3$  nano composite increased. The added silver nanoparticles occupied lattice vacancies and  $\gamma$ -phase for alumina didn't change. Determined by Scherer's equation (Eq. (1)) method from the XRD half-peak broadening

$$D = \left( \frac{0.9\lambda}{\beta \cos\theta} \right) \quad (\text{Eq. (1)})$$

where D is the crystallites size in nm,  $\lambda$  the radiation of wavelength (Cu  $K\alpha_1$ , 0.15406 nm),  $\beta$  the full width at half of the maximum in radians and  $\theta$  is the Bragg-angle. The results are shown in Table 1.

Table 1. Crystal size of Ag and  $\text{Al}_2\text{O}_3$  particles by Scherer's equation

Sample	Crystallite size of Ag (nm)	Crystallite size of $\text{Al}_2\text{O}_3$ (nm)
a	-	1.75
b	3.07	1.99
c	3.74	3.23
d	4.99	4.54

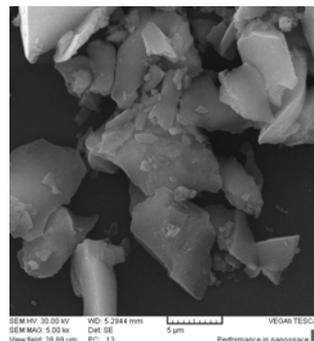
The results of BET and BJH measurements for nanocomposites are shown in Table 2. The surface area of the  $\text{Ag}\backslash\text{Al}_2\text{O}_3$  nanocomposite gradually decreased with increasing silver crystallites and protected Alumina phase. Catalysis with this property is very much useful for industrial applications specifically in petroleum industry. Also with increasing silver average pore volume slightly reduced but mean pore diameter increased indicating that combining with our XRD spectra (particle size of silver), silver nanoparticles are distributed in pores of alumina surfaces. BJH analysis showed average pore volumes in samples

b, c and d smaller than that of the (sample a). (Sample a) has microporous pores and others samples have mesoporous pores. In mesoporous materials due to larger pores more number of molecules interact with each other, and they show better catalytic properties [23]. Hysteresis loops occur at low relative pressures (around 0.5) suggesting a narrow pore size distribution [24]. The pore size distribution confirms this assertion. Well-dispersed to solution tends to produce small pore size and homogeneous product.

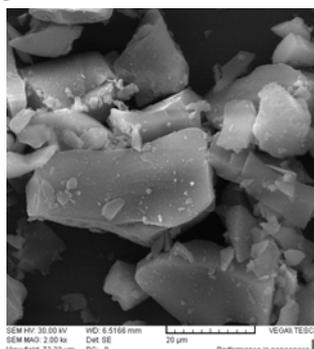
The microstructure of the sample is analyzed by using SEM and TEM microscopy. SEM micrograph shown with the further increase of silver loading there are smaller bulk, this indicating interaction the surface charge because smaller particle with more number of particles make a larger bulk and larger surface area, but unlike larger particles with lower number of particles make a smaller bulk and smaller surface area. This result is in agreement with the value obtained from XRD and BET analysis. TEM microscopy (fig3) shows Ag particles were mostly around 3.54 nm in size and clearly shows that the Ag particles are almost uniformly dispersed on the  $\text{Ag}/\text{Al}_2\text{O}_3$  nanocomposite [25]. The SEM, BET, BJH and analysis of samples showed that silver is distributed quite homogeneously in pores of Alumina surface, no changes was seen surface of Alumina.

**Table 2:** results of BET and BJH measurements

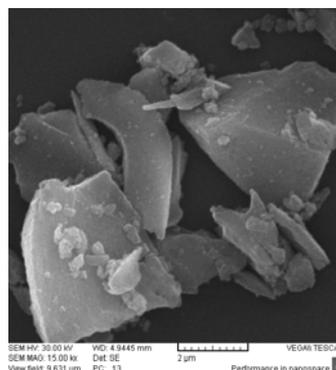
Sample	specific surface area( $\text{m}^2/\text{g}$ )	Average pore volume( $\text{cm}^3/\text{g}$ )	Mean pore diameter(nm)
a	392.79	0.303	3.09
b	148.08	0.209	5.64
c	125.15	0.218	6.98
d	102.25	0.209	8.19



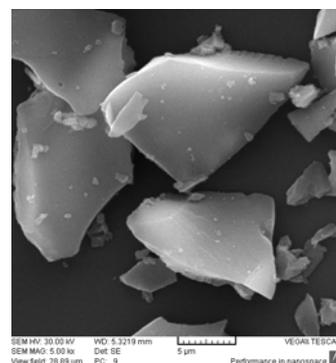
a) 1wt.% $\text{Ag}/\text{Al}_2\text{O}_3$



b) 5wt.% $\text{Ag}/\text{Al}_2\text{O}_3$

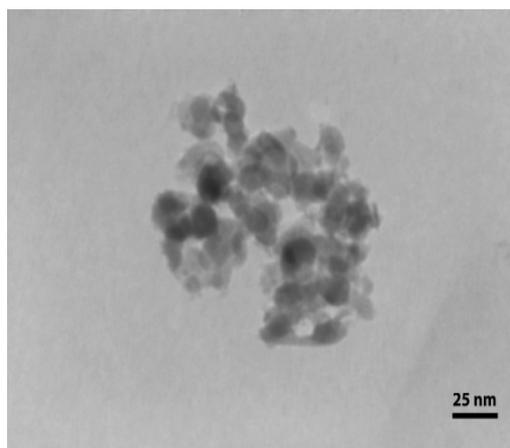


b) 10wt.% $\text{Ag}/\text{Al}_2\text{O}_3$



d) $\text{Al}_2\text{O}_3$

**Fig. 2.** (a),(b),(c)and (d) The SEM micrographs of  $\text{Ag}/\text{Al}_2\text{O}_3$  nanocomposite and  $\text{Al}_2\text{O}_3$  nanoparticles

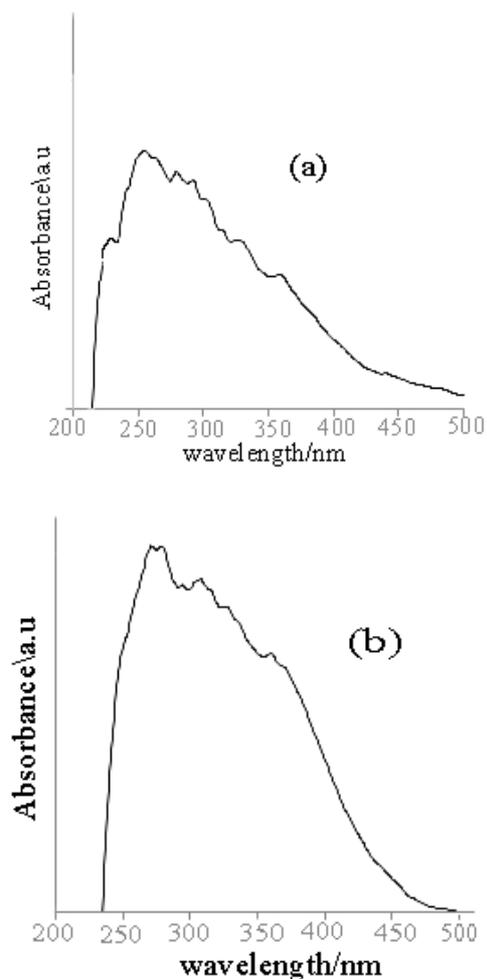
10wt. %Ag/Al<sub>2</sub>O<sub>3</sub>

**Fig. 3.** Shows TEM micrograph of 10wt. %Ag/Al<sub>2</sub>O<sub>3</sub> nanocomposite

UV/Vis was taken to discover the details of the state of the supported Ag. Fig. 8 shows the spectra of (a) the 5wt. %Ag/Al<sub>2</sub>O<sub>3</sub> that synthesized in 400°C and (b) the 5wt. % Ag/Al<sub>2</sub>O<sub>3</sub> that synthesized in 500°C. UV/Vis analysis demonstrates the existence of different states of the Ag species because the presence Ag<sup>+</sup> species makes the high catalytic activity of Ag-Al<sub>2</sub>O<sub>3</sub>.

The band at 215 nm is attributed to the 4d<sup>10</sup> to 4d<sup>9</sup>5s<sup>1</sup> transition of highly dispersed Ag<sup>+</sup> ions [26, 27], 275 nm is assigned to small Ag<sup>δ+</sup><sub>n</sub> clusters [28–29] and the bands at 350 and 450 nm are attributed to metallic silver (Ag<sup>0</sup>) particles [26–29]. Sato et al. confirmed an absorption band of metallic Ag<sub>n</sub> clusters at 352 nm for 4% Ag/Al<sub>2</sub>O<sub>3</sub> [28]. Bogdanchikova et al. also reported an absorption band of Ag<sup>0</sup> at a wavelength > 390 nm for reduced Ag/Al<sub>2</sub>O<sub>3</sub> [27]. Accordingly, we can estimate that the bands at 350 and 450 nm are due to metallic Ag<sup>0</sup> particles. The absorption band around 276 nm and 350 ~ 400 nm was observed in both samples which respectively was assigned to Ag<sup>δ+</sup><sub>n</sub> clusters and metallic silver (Ag<sup>0</sup>) particles. In table (b) The band at 215 nm is attributed to Ag<sup>+</sup>

ions, but in table (a) this band isn't observed. Ag/γ-Al<sub>2</sub>O<sub>3</sub> catalyst, highly dispersed Ag<sup>+</sup> ions were required for HC-SCR [30, 31].



**Fig. 4.** Shows the spectra of (a) the 5wt. %Ag/Al<sub>2</sub>O<sub>3</sub> that synthesized in 400 °C and (b) the 5wt. % Ag/Al<sub>2</sub>O<sub>3</sub> that synthesized in 500 °C

The high activity of Ag-Al<sub>2</sub>O<sub>3</sub> for SCR of NO to N<sub>2</sub> by n-octane and higher hydrocarbons was believed to be due to the high dispersion of the Ag species and Ag<sup>+</sup> ions [32]. Large amount of active Ag<sup>+</sup> ions and Ag<sup>δ+</sup><sub>n</sub> clusters in Ag/γ-Al<sub>2</sub>O<sub>3</sub> showed a uniform distribution of silver species on alumina support and an excellent catalytic behaviour for NO<sub>x</sub> decomposition without any N<sub>2</sub>O emission [18].

#### 4. Conclusion

Ag/Al<sub>2</sub>O<sub>3</sub> Nanocomposite have been successfully synthesized by solution combustion technique. Silver nanoparticles with different amounts (1wt%, 5wt% and 10wt%) were added to alumina, With increasing amount of Ag, the mean size of the Ag and Al<sub>2</sub>O<sub>3</sub> particles in the Ag/Al<sub>2</sub>O<sub>3</sub> nano composite increased and surface area of the Ag/Al<sub>2</sub>O<sub>3</sub> nanocomposite gradually decreased but mean pore diameter increased while the pore volume was not changed appreciably for this composite, its shows The added silver nanoparticles occupied lattice vacancies and  $\gamma$ -phase for alumina didn't change. Ag particles were mostly around 3.54 nm in size and clearly shows that the Ag particles are almost uniformly dispersed on the Ag/Al<sub>2</sub>O<sub>3</sub> nanocomposite. Analysis demonstrates the existence of different states of the Ag species, and increasing synthesis temperature emitted Ag<sup>+</sup> species as a result synthesized sample in 400C<sup>0</sup> have good potential to be used as catalytic with high activities.

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